

## Chemistry Of Interhalogen Compounds

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### Chemistry Of Interhalogen Compounds

The Interhalogen Compounds . Interhalogen compounds are formed by reactions between different halogens. All possible interhalogen compounds of the type XY are known. Bromine reacts with chlorine, for example, to give BrCl, which is a gas at room temperature.

### The Chemistry of the Halogens - Purdue University

Astatine is known to react with its lighter homologs iodine, bromine, and chlorine in the vapor state; these reactions produce diatomic interhalogen compounds with formulas AtI, AtBr, and AtCl. The first two compounds may also be produced in water – astatine reacts with iodine/ iodide solution to form AtI, whereas AtBr requires (aside from ...

## **Astatine - Wikipedia**

CHM101: Chemistry and Global Awareness (Gordon) ... To know what types of elements bond to form covalent compounds. ... Chlorine monofluoride is a volatile interhalogen compound. Answer b  $\text{HI}$  is used in organic and inorganic synthesis as one of the primary sources of iodine.

## **4.8: Covalent Bonding and Formula Writing - Chemistry ...**

Iodine is a chemical element with the symbol I and atomic number 53. The heaviest of the stable halogens, it exists as a lustrous, purple-black non-metallic solid at standard conditions that melts to form a deep violet liquid at 114 degrees Celsius, and boils to a violet gas at 184 degrees Celsius. The element was discovered by the French chemist Bernard Courtois in 1811, and was named two ...

## **Iodine - Wikipedia**

SRMJEEE 2021- SRM University will be conducting SRMJEEE 2021 exam in two phases as a remote proctored test. The first phase of SRMJEEE 2021 is scheduled to be conducted from May 23 to 24 while phase 2 exam will be conducted from July 25 to 26. the exam conducting authority has released the SRMJEEE 2021 application form on the official website (srmist.edu.in) on December 12.

## **SRMJEEE 2021: Application Form (OUT), Exam Dates ...**

(iii) Fluorine forms the largest number of interhalogen compounds amongst the halogens. Answer:  
(i)  $\text{NH}_3$  acts as a good ligand due to presence of lone pair of electron which it can readily donate.  
(ii) Refer Ans. to Q.75. (iii) It is because fluorine is the most electronegative and strongest oxidising agent. Question 114: Account for the following:

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